

**SYNTHESIS AND SELF-ASSEMBLY STUDIES
OF GLYCOSIDE SURFACTANTS AND CHROMONICS**

FARAMARZ ALIASGHARI SANI

**THESIS SUBMITTED IN FULFILLMENT OF THE
REQUIREMENTS FOR THE DEGREE OF DOCTOR OF
PHILOSOPHY**

**DEPARTMENT OF CHEMISTRY
FACULTY OF SCIENCE
UNIVERSITY OF MALAYA
KUALA LUMPUR**

2012

UNIVERSITI MALAYA
ORIGINAL LITERARY WORK DECLARATION

Name of Candidate: **FARAMARZ ALIASGHARI SANI** (I.C/Passport No: **K17141517-New** (G2236394-Old))

Registration/Matric No: **SHC080065**

Name of Degree: **Doctor of Philosophy**

Title of Thesis ("this Work"):

SYNTHESIS AND SELF-ASSEMBLY STUDIES OF GLYCOSIDE SURFACTANTS AND CHROMONICS

Field of Study: **Organic Chemistry Including Liquid Crystal and Nanotechnology**

I do solemnly and sincerely declare that:

- (1) I am the sole author/writer of this Work;
- (2) This Work is original;
- (3) Any use of any work in which copyright exists was done by way of fair dealing and for permitted purposes and any excerpt or extract from, or reference to or reproduction of any copyright work has been disclosed expressly and sufficiently and the title of the Work and its authorship have been acknowledged in this Work;
- (4) I do not have any actual knowledge nor do I ought reasonably to know that the making of this work constitutes an infringement of any copyright work;
- (5) I hereby assign all and every rights in the copyright to this Work to the University of Malaya ("UM"), who henceforth shall be owner of the copyright in this Work and that any reproduction or use in any form or by any means whatsoever is prohibited without the written consent of UM having been first had and obtained;
- (6) I am fully aware that if in the course of making this Work I have infringed any copyright whether intentionally or otherwise, I may be subject to legal action or any other action as may be determined by UM.

Candidate's Signature
Date

Subscribed and solemnly declared before,

Witness's Signature
Date

Name:

Designation:

Abstract

This thesis focused on the synthesis and self assembly studies of compounds involving carbohydrates. Two different types of materials were studied, *i.e.* surfactants and chromonics. Physical investigations of the compounds were conducted by TGA, DSC, OPM, UV-Vis spectroscopy, ^1H NMR and surface tension measurements. Sugar-based surfactants are interesting compounds for pharmaceutical and personal care products. Current commercially available surfactants such as alkyl poly glucoside surfactants (APGs) are prepared from low miscibility of sugars and fatty alcohols. In order to solve the miscibility problems of the starting materials, homogenizers are required. This, however, leads to impurities in products, which affects the use of the surfactant for life science applications. Therefore, this work focused on an economic preparation of pure glycoside surfactants. The synthesis approach applied a separation of glycosides and a coupling of sugar and fatty alcohols with different length and chain branching by click chemistry. Twenty one alkyl triazole glycoside surfactants (ATGs) were prepared with more than 80% yield. Twelve of these were anomeric pure products (>95% purity) and nine were technical products of α/β anomeric mixtures. The materials were characterized for their liquid crystal behaviours. Contact penetration studies showed a whole range of lyotropic phases from lamellar to cubic and hexagonal. ATGs' CMCs were found to be lower than those of APG surfactants. An increase in their chain length meant a decrease in CMC values. Hence these materials could be identified for oil-based surfactants applications.

Chromonics or lyotropic chromonic liquid crystals (LCLCs) are formed by the self-association of aromatic disk-shaped molecules with hydrophilic groups at the periphery in aqueous solutions. The chromonics are assembled from π - π interactions of the aromatic cores. This leads to aggregates based on stacking of the molecules. Most chromonic molecules are based on ionic structures. The research embraced the synthesis and assembly studies of non ionic chromonics consisting of triphenylene-based units surrounded by glycosides. The key point for the synthesis of triphenylene core was oxidative trimerization of veratrole and guaiacol under anhydrous conditions in the presence of ferric chloride. A symmetric compound with six sugars and an asymmetric one with three sugar units were synthesized. The materials were of purity over 95% and more than 75% yield. Due to the anisotropic effect on the aromatic ring on ^1H NMR, the chemical shift on the aromatic ring changed when the concentration was increased. This means the materials formed aggregation and enabled the determination of critical aggregation concentrations (CAC). Moreover, this property also showed temperature dependency. At higher concentrations and under examination by polarizing light microscopy, the chromonic exhibited liquid crystalline properties (Col phase).

Abstrak

Tesis ini memberi fokus terhadap kajian sintesis dan penyusunan diri yang melibatkan sebatian karbohidrat. Dua jenis bahan telah dikaji, iaitu surfaktan dan kromonik. Kajian fizikal sebatian telah dijalankan menggunakan TGA, DSC, OPM, spektroskopi UV-Vis, ^1H NMR dan pengukuran ketegangan permukaan. Surfaktan berasaskan gula merupakan sebatian yang menarik untuk dijadikan produk farmaseutikal dan penjagaan peribadi. Surfaktan komersial semasa seperti surfaktan alkil poli glukosida (APGs) telah disediakan daripada gula dan alkohol lemak yang rendah kebolehcampuran. *Homogenizers* diperlukan untuk menangani masalah kebolehcampuran bahan permulaan. Walau bagaimanapun, hal ini telah menyumbang kepada bendasing di dalam produk dan seterusnya memberi kesan terhadap penggunaan surfaktan bagi aplikasi sains hayat. Kerja ini tertumpu kepada penghasilan surfaktan glikosida tulen secara ekonomi. Pendekatan sintesis melibatkan pemisahan glikosida dan padanan gula serta alkohol lemak dengan pelbagai kepanjangan dan rantaian cabang melalui kimia klik. Sebanyak 21 surfaktan alkil triazole glikosida (ATGs) telah disediakan dengan peratusan hasil melebihi 80%. Dua belas daripadanya merupakan produk anomer tulen ($>95\%$) dan sembilan yang lain adalah produk teknikal yang terdiri daripada campuran anomer α/β . Sifat hablur cecair bahan-bahan ini telah dikaji. Kajian ‘contact penetration’ telah menunjukkan kepelbagaian fasa liotropik dari lamelar kepada kubik dan heksagon. Kepekatan kritikal miselar (CMC) bagi ATG adalah lebih rendah daripada surfaktan APG. Peningkatan panjang rantaian mereka, mengurangkan nilai CMC. Oleh itu, bahan ini berkemungkinan boleh digunakan untuk aplikasi surfaktan berasaskan minyak.

Kromonik atau hablur cecair kromonik liotropik (LCLCs) terbentuk dari penyatuan diri molekul-molekul berbentuk cakera aromatik dengan kumpulan hidrofilik di pinggir larutan akueus. Kromonik ini menyusun disebabkan oleh interaksi π - π teras aromatik. Ini membawa kepada pengagregatan berdasarkan kepada penyusunan molekul. Kebanyakan molekul kromonik berdasarkan struktur ionik. Penyelidikan ini melibatkan kajian sintesis dan penyusunan diri kromonik bukan-ionik yang terdiri daripada unit berdasarkan triphenylene yang dikelilingi oleh glikosida. Tumpuan utama bagi sintesis teras triphenylene adalah oksidatif *trimerization veratrole* dan guaiacol di bawah keadaan yang kontang dengan kehadiran ferik klorida. Satu sebatian simetri dengan enam unit gula dan satu sebatian tidak-simetri dengan tiga unit gula telah disintesis. Bahan-bahan ini telah dihasilkan dengan ketulenan yang tinggi melebih 95% dan peratusan hasil lebih daripada 75%. Disebabkan oleh kesan tidak-isotropik pada gelang aromatik dalam ^1H NMR, anjakan kimia pada gelang aromatik telah berubah dengan peningkatan kepekatan. Ini bermakna, bahan-bahan telah membentuk pengagregatan dan membolehkan penentuan kepekatan kritikal pengagregatan (CAC). Tambahan pula, sifat ini juga menunjukkan kebergantungan terhadap suhu. Pada kepekatan yang lebih tinggi dan pemerhatian di bawah mikroskop polarisasi cahaya, kromonik ini menunjukkan sifat hablur cecair (fasa Kol).

Acknowledgements

I wish to express my sincere gratitude to all people who have contributed to the realization of this research.

I would like to express my deepest regards and sincere gratitude to the head of the glycolipid and science technology group, **Professor Dr. Rauzah Hashim**, for her guidance, enthusiasm and support. I really appreciate her constant encouragement throughout my Ph.D. programm and her ready availability on all matters. Her kindness has encouraged me to overcome the difficulties not only in scientific work but also in social life.

I am deeply indebted to **Associate Professor Dr. Thorsten Heidelberg** for his countless advice, guidance and enthusiastic supervision throughout years. He helped me to understand the meaning of carbohydrate chemistry. I would also like to thank for his assistance in completing the dissertations.

I would like to thank my parents, Mother “**Marziyeh**”, Father “**Esmaeil**” , my brother “**Fariborz**”, and two sisters “**Farinaz & Faranak**” for all the times I needed support while achieving my PhD.

I would like to express my deepest grateful to my wife “**Narges**” for her patient and providing a perfect environment to study.

I especially appreciate all my lab mates, past and present, for helpful discussions and providing a friendly environment in which to work and study.

I would like to express my deepest thanks to the **Managing Director (CEO)** **Eng. Seyyed Ahmad Kakhki** (previous), **Dr. Jahangir Mallaki** (present) and **Authorized Pharmacist, Strategic Studies, Management Officer and R&D Director Dr. Ali Molavi** all of Daroo Sazi Samen Co. (**Samen Pharmaceutical Company**) for their support on my PhD. programme.

I am also grateful to the manager of QC. labs **Dr. Seyyed Alireza Kalati** and regulatory affairs officer **MSc. Mehrdad Taghavi Gilani** and other colleagues as well as members of Daroo Sazi Samen Co.

The financial supports provided by the **University of Malaya**, Faculty of Science and the Department of Chemistry are gratefully appreciated, I also appreciate the research assistantship under the High Impact Research grant from the University of Malaya and the Ministry of the Higher education (MOHE), UM.C/625/1/HIR/MOHE/05.

Finally, I want to dedicate this thesis to **my parents and my wife** for their love and understanding in every step of my life. To my beloved son “**Shahrad**”, his smile has the greatest meaning to me.

Thank you very much one and all.

Table of Contents

List of Figures	xv
List of Tables	xxii
List of Abbreviations	xxiii
1. Introduction	1
1.1. Non ionic surfactants	1
1.2. Novel chromonic compounds	3
1.3. Outline of later chapters	4
1.4. Objective of the study	4
2. Literature review	6
2.1. Introduction	6
2.2. Amphiphilic	6
2.2. Glycolipid	6
2.3. Structure of surfactants	8
2.4. Different classes of surfactants	11
2.4.1. Anionic surfactants	11
2.4.2. Cationic surfactants	11
2.4.3. Zwitterionic surfactants	12
2.4.4. Non ionic surfactant	13
2.5. Behaviour of surfactants in solutions	14
2.6. Micelle formation of surfactant in solutions	15
2.7. Aggregation behaviour of surfactants	16
2.8. Lyotropic liquid crystals	16
2.9. Krafft point and cloud point	19
2.10. Hydrophilic-Lipophilic Balance	20
2.11. Thermotropic liquid crystals	21
2.12. Surfactants in industry	21
2.13. Biodegradability	22
2.14. Use of sugar-based surfactants	23

2.15.	Types of sugar-based surfactants	23
2.16.	Synthesis of glycolipids	25
2.16.1.	Glycosylation strategy	25
2.16.2.	Synthesis of glycosides	26
2.16.3.	Fisher glycosidation synthesis	26
2.16.4.	Neighbor group participation	27
2.16.5.	Synthesis alkyl polyglucoside (APG)	28
2.17.	Click chemistry	29
2.18.	Generation of natural product analog by click chemistry	31
2.19.	Drug discovery approaches based on click chemistry	31
2.20.	Liquid crystals	32
2.21.	Calamitic mesophases	34
2.22.	Discotic mesophases	34
2.23.	Applications of liquid crystal	37
2.24.	Supermolecular discotics	37
2.25.	Polycyclic aromatic hydrocarbons	39
2.26.	Synthesis of PAHs	39
2.27.	Chromonic liquid crystal system	45
2.28.	LCLCs for detection of biological application	47
2.29.	Important factors affecting chromonic aggregation	48
3.	Materials and methods	50
3.1.	General methods	50
3.1.1.	NMR spectroscopy	50
3.1.2.	High resolution mass spectrometry	50
3.1.3.	UV-Vis spectroscopy	50
3.1.4.	Fluorescence spectroscopy	50
3.1.5.	Fourier transforms infrared spectroscopy	50
3.1.6.	Elemental analysis	51
3.1.7.	Differential scanning calorimetry (DSC)	51
3.1.8.	Thermo gravimetric analysis (TGA)	51

3.1.9.	Optical polarizing microscopy (OPM)	51
3.1.10.	Surface tension measurement	51
3.1.11.	Polarimeter	51
3.1.12.	Other facilities	52
3.2.	Materials	52
3.2.1.	Chemicals and solvents	52
3.2.2.	Chromatography	52
3.2.3.	Synthetic procedure	52
3.2.3.1.	Part A: Synthesis of surfactants (alkyl triazole surfactants)	52
3.2.3.1.1.	General procedure of glycosidation: Synthesis of (α/β) propargyl-glucoside	51
3.2.3.1.2.	General glycosidation: Synthesis of propargyl 2,3,4,6-tetra-O-acetyl- β -D-glucopyranoside	53
3.2.3.1.3.	General deacetylation procedure (surfactants)	53
3.2.3.1.4.	General bromination procedure: Synthesis of alkyl bromide	54
3.2.3.1.5.	General tosylation procedure: Synthesis of (Z)-9-octadecenyl 4-tolyl sulfonate	54
3.2.3.1.6.	General chlorination procedure: Synthesis of 1-Chloro-2-hexyl-decane	55
3.2.3.1.7.	General azidation: Synthesis of alkyl azide	55
3.2.3.1.8.	Synthesis of α -propargyl glucoside tetraacetate	55
3.2.3.1.9.	General CLICK chemistry procedure: Synthesis of pure products	56
3.2.3.1.10.	General CLICK chemistry procedure: Synthesis of technical products	56
3.2.3.2.	Part B: Synthesis of Chromonic	57
3.2.3.2.1.	Synthesis of 2,3,6,7,10,11-hexamethoxytri- phenylene (HMTP)	57
3.2.3.2.2.	Synthesis of 2,3,6,7,10,11-hexahydroxytri- phenylene (HHTP)	57
3.2.3.2.3.	Synthesis of 2-bromoethyl 2,3,4,6-tetra-O- acetyl- β -D-glucopyranoside	57
3.2.3.2.4.	Synthesis of 2,6,11-trihydroxy-3,7,10-tri- methoxytriphenylene	58

3.2.3.2.5. Synthesis of 2,3,6,7,10,11-hexakis-[2-(2,3,4,6-tetra- <i>O</i> -acetyl- β -D-glucopyranosyl oxyethyl]-triphenylene	58
3.2.3.2.6. Synthesis of 2,6,11-tri-(ethyl 2,3,4,6-tetra- <i>O</i> -acetyl- β -D-glucopyranoside) 3,7,10-tri-phenylene	58
3.2.3.2.7. General deacetylation procedure for triphenylene derivatives	59
3.2.3.3. Chemical product analysis	59
3.2.3.3.1. Nuclear magnetic resonance (NMR)	59
3.2.3.3.2. Solvent selection	61
3.2.3.3.3. FT-IR spectroscopy	63
3.2.3.3.4. Attenuated total reflectance (ATR) spectroscopy	64
3.2.3.3.5. High resolution mass spectrometry	65
3.2.3.3.5. UV-Vis spectroscopy	66
3.2.3.3.6. Fluorescence spectroscopy	66
3.2.3.3.7. Elemental analysis	66
3.2.3.3.8. Polarimeter	67
3.2.3.3.9. Differential scanning calorimetry (DSC)	67
3.2.3.3.10. Thermo gravimetric analysis (TGA)	68
3.2.3.3.11. Optical polarizing microscopy (OPM)	68
3.2.3.3.12. Surface tension measurement	69
3.2.3.3.13. Krafft point and cloud point measurements	69
4. Results and discussion	71
4.1. ATGs (alkyl triazole glucoside surfactant)	71
4.1.1. Synthesis	71
4.1.1.1. Synthesis of propargyl glycoside	71
4.1.1.2. Mechanism of Fischer glycosidation of propargyl glycoside	72
4.1.1.3. Anomeric effects and propargyl glycoside	72
4.1.1.4. Synthesis of propargyl 2,3,4,6-tetra- <i>O</i> -acetyl- α -D-glucopyranoside	73

4.1.1.5.	Synthesis of propargyl 2,3,4,6-tetra- <i>O</i> -acetyl- β -D-glucopyranoside	73
4.1.1.6.	Synthesis of guerbet bromides	74
4.1.1.7.	Synthesis of guerbet chloride	74
4.1.1.8.	Oleyl tosylates ((Z)-9-octadecenyl 4- toluyl-sulfonate)	76
4.1.1.9.	Synthesis of alkyl azide	76
4.1.1.10.	Development of new surfactants by the reaction of CLICK chemistry	77
4.1.1.11.	The mechanism approach	78
4.1.1.12.	Technical alkyl triazole glycoside surfactants	81
4.1.1.13.	β -alkyl triazole glycoside surfactants	84
4.1.1.14.	α -alkyl triazole glycoside surfactants	87
4.1.2.	Physicochemical properties	88
4.1.3.	Liquid crystal behaviour	93
4.1.3.1.	Thermotropic Properties	93
4.1.3.2.	Lyotropic properties	97
4.2.	Chromonics (Triphenylen Glycosides)	99
4.2.1.	Synthesis	99
4.2.1.1.	Synthesis of 2-bromoethyl 2,3,4,6-tetra- <i>O</i> -acetyl- β -D-glucopyranoside	99
4.2.1.2.	Synthesis of hexa-methoxytriphenylene	99
4.2.1.3.	Synthesis of 2,3,6,7,10,11-hexahydroxytriphenylene (HHTP)	102
4.2.1.4.	Synthesis of 2,6,11-trihydroxy-3,7,10-trimethoxytri-phenylene (asym-TP(OH) ₃ -(OCH ₃) ₃)	103
4.2.1.5.	Synthesis of 2,3,6,7,10,11-hexakis-(2-[2,3,4,6-tetra- <i>O</i> -octyl- β -D-glyco-pyranosyl-oxy]-ethyl-oxy)-tri-phenylene	105
4.2.1.6.	Synthesis 2,6,11-trikis-(2-[2,3,4,6-tetra- <i>O</i> -octyl- β -D-glycopyranosyl-oxy]-ethoxy)-3,7,10-methoxy-triphenylene	106
4.2.1.7.	General procedure for the De- <i>O</i> -Alkylation	106
4.2.2.	Thermal behaviour	111
4.2.3.	IR study	115

4.2.4.	UV-Vis study	116
4.2.5.	Self-assembly chromonics in water	117
4.2.5.1.	NMR investigation	117
4.2.5.2.	UV-Vis spectroscopy	120
4.2.5.3.	ATR-FTIR data analysis	120
5. Conclusion		122
5.1.	Alkyl triazole glycoside surfactants (ATGs)	122
5.2.	Chromonics (LCLCs)	123
5.3.	Comparison of the assembly behaviour for ATGs and LCLCs	123
6. Appendices		125
7. References		172

List of Figures

Chapter 1

Figure 1.1. Consumption of surfactants in Europe (adapted based on ref. (Greindl, 2011))	2
---	---

Chapter 2

Figure 2.1. A schematic model of the Gram-negative cell envelope present in <i>E. coli</i> (image from (Anthony, 2009))	7
Figure 2.2. Diagram of ABO blood groups (image from (Invicta, 2006))	7
Figure 2.3. Schematic representation of a surfactant	8
Figure 2.4. Surfactant dodecyl chain with different head group	9
Figure 2.5. Surfactants with different hydrophobic moieties	9
Figure 2.6. Different kind of surfactants	10
Figure 2.7. An anionic surfactant	10
Figure 2.8. Examples of cationic surfactants	11
Figure 2.9. Examples of preparation of cationic surfactants	11
Figure 2.10. Examples of zwitterionic surfactants	12
Figure 2.11. Examples of non ionic surfactants	12
Figure 2.12. Scheme illustrates water molecules at liquid-air interface	13
Figure 2.13. Equilibrium between surfactants solution at low concentration in water	13
Figure 2.14. Arrangement of surfactants in a solution	14
Figure 2.15. Illustration of various micellar morphologies (Tsujii, 1997)	15
Figure 2.16. General surfactant liquid crystalline phases	16
Figure 2.17. Phase behaviour of amphiphilic molecules in water (Diagram modified based on ref. (Tiddy, 1980))	17
Figure 2.18. Schematic phase diagram for a surfactant (adapted and modified from (Wang, 2002))	18
Figure 2.19. Structures of some poloyl surfactants	22

Figure 2.20. Structures of sucrose ester and general lactose, lactitol mono-ester surfactant	24
Figure 2.21. General glycosylation reaction	25
Figure 2.22. Preparation of oxacarbenium ion intermediate	25
Figure 2.23. Fisher glycosidation	26
Figure 2.24. Effect of neighbor group	27
Figure 2.25. Illustration of anomeric effect	28
Figure 2.26. 1,2,3-Triazole formation via Huisgen 1,3-dipolar cycloaddition	29
Figure 2.27. Thermal and Cu(I)-catalyzed 1,3-dipolar cycloaddition.	30
Figure 2.28 Generation of vancomycin analogs	31
Figure 2.2.1. Structure and phase characterizations of cholesteryl benzoate	32
Figure 2.2.2. Example of calamitic (rod like) liquid crystal	33
Figure 2.2.3. Example of disk-like liquid crystal	33
Figure 2.2.4. Schematic representation of nematic phase	34
Figure 2.2.5. Schematic representation of smetic A (S_mA) phase	34
Figure 2.2.6. Explanation of benzoate derivative mesophases	35
Figure 2.2.7. Molecular structures commonly used in the preparation of discotic mesogens	35
Figure 2.2.8. Molecular structure for phthalocyanine	36
Figure 2.2.9. Schematic representation of columnar phases	37
Figure 2.2.10. Formation of 'discogenic' mesophase by non-covalent Interactions	38
Figure 2.2.11. Layer structure of graphite (adapted from ref. (Clar, 1972))	40
Figure 2.2.12. Example of the Harworth synthesis	40
Figure 2.2.13. Example of Diels-Alder cycloaddition for construction of PAHs	40
Figure 2.2.14. Muller's synthesis of the rhombus-shaped PAH	41

Figure 2.2.15. Photochemical cyclization to obtain PAHs	41
Figure 2.2.16. Synthesis of PAHs by FVP method	42
Figure 2.2.17. Synthesis of PAHs by route involving extraction of sulfur	42
Figure 2.2.18. Synthesis of Kekulene according to the Diederich and Staab method	43
Figure 2.2.19. Cyclodehydrogenation of hexa-phenylbenzene to HBC	43
Figure 2.2.20. Synthesis of triphenylene	44
Figure 2.2.21. Synthesis of triphenylene derivatives	44
Figure 2.2.22. Structures for two molecules that form chromonic liquid crystals, DSCG and SSY	45
Figure 2.2.23. Schematic diagram of LCLC aggregates in (a) I phase, (b) N phase and (c) C (M) phase	46
Figure 2.2.24. The scheme of lyotropic chromonic LCs biosensor modified based on (Shiyanovskii, 2005)	48

Chapter 3

Figure 3.1. Example of a superconducting NMR magnet (Adapted and re-drawn from reference (Richards, 2011))	60
Figure 3.2. Illustration of the “edge effect” problems (adapted and re-drawn from reference (Richards, 2011))	60
Figure 3.3. Shows a method for filtering NMR solution and undissolved material in solution (adapted and re-drawn from reference (Richards, 2011))	61
Figure 3.4. NMR spectrum of a propargyl glycoside	62
Figure 3.5. Representation of schematic example of HMQC spectrum	63
Figure 3.6. Concept of ATR spectroscopy (image from (PerkinElmer, 2005))	64

Chapter 4

Figure 4.1.1. Fischer glycosidation	71
Figure 4.1.2. Illustrates the propargyl glycosidation	71

Figure 4.1.3. Mechanism of Fischer glycosidation	72
Figure 4.1.4. Illustration of anomeric effect	73
Figure 4.1.5. Mechanism of glycosidation of protected sugar	74
Figure 4.1.6. Mechanism of bromination of guerbet alcohol	75
Figure 4.1.7. Mechanism of chlorination of guerbet alcohol	75
Figure 4.1.8. Mechanism of tosylation	76
Figure 4.1.9. Scheme of azidation	76
Figure 4.1.10. IR spectra of C ₁₈ H ₃₇ N ₃ and C ₁₈ H ₃₈ Br	77
Figure 4.1.11. Example of click chemistry	77
Figure 4.1.12. Classical thermal cycloaddition reaction	78
Figure 4.1.13. Effect of copper catalysis in 1,3-cycloaddition	78
Figure 4.1.14. Proposed reaction mechanism of click chemistry (adapted from (Himo, 2004))	79
Figure 4.1.15. Illustrates the reaction of copper(I) acetylides with organic azides, determined by DFT calculation. (adopted from (Himo <i>et al.</i> , 2004))	79
Figure 4.1.16. Coupling reaction by click chemistry	81
Figure 4.1.17. ¹ H NMR and ¹³ C NMR spectra of technical alkyl triazole surfactant	83
Figure 4.1.18. Schematic syntheses of pure β anomer peracetylated ATGs surfactants	84
Figure 4.1.19. HMQC assignment	85
Figure 4.1.20. Synthesis route of β -anomer peracetylated ATGs surfactants	86
Figure 4.1.21. Example of two ATGs products and their concentrated solutions	89
Figure 4.1.22. Curve of surface tension values and CMC for β -ATG-C ₁₀	90
Figure 4.1.23. Comparison curves of surface tension of technical, α - and β -anomer of ATG-C ₁₂	90
Figure 4.1.24. Birefringent showed by (left) α -ATG-C ₈ and (right) β -ATG-C ₁₀ at 25°C	93

Figure 4.1.25. Examples of thermotropic behaviour of ATGs (50 x magnification)	94
Figure 4.1.26. Example of DSC study on β -ATG-C ₁₀	96
Figure 4.1.27. Effect of length chain on DSC measurements on β -ATG products	97
Figure 4.1.28. Contact penetration of β -ATG-C ₁₀	98
Figure 4.2.1. Synthesis route of the 2-bromoethyl 2,3,4,6-tetra- <i>O</i> - acetyl- β -D-gluco-pyranoside	99
Figure 4.2.2. Oxidative trimerization of veratrole to HMTP	99
Figure 4.2.3. Mechanism proposed for trimerization of veratrole	100
Figure 4.2.4. Structure of hexamethoxy triphenylene (HMTP)	101
Figure 4.2.5. Dealkylation of hexamethoxy triphenylene	102
Figure 4.2.6. Structure of hexahydroxy triphenylene (HHTP)	103
Figure 4.2.7. Formation of the symmetric and asymmetric structure isomers	103
Figure 4.2.8. ¹ H NMR spectra assignment of structural isomer of triphenylene	104
Figure 4.2.9. ¹ H NMR spectra assignment of structural triphenylene	104
Figure 4.2.10. Synthetic scheme for hexakis-(2-[2,3,4,6-tetra- <i>O</i> -octyl- β -D-glycopyranosyl-oxy]-ethoxy)-triphenylene	105
Figure 4.2.11. Synthetic scheme for 2,6,11-tris-(2-[2,3,4,6-tetra- <i>O</i> -octyl- β -D-glycopyranosyl-oxy]-ethoxy-3,7,10-trimethoxy triphenylene	106
Figure 4.2.12. Synthetic scheme for hexakis-[2-(β -D-glycopyranosyl-oxy)-ethoxy]-triphenylene, (GlcOC ₂ H ₄ O) ₆ TP	107
Figure 4.2.13. DEPT 135 and ¹³ C NMR spectra for (GlcOC ₂ H ₄ O) ₆ TP	107
Figure 4.2.14. HMQC spectrum for (GlcOC ₂ H ₄ O) ₆ TP	108
Figure 4.2.15. Expansion of the ¹ H NMR spectra of (GlcOC ₂ H ₄ O) ₆ TP and asym-(GlcOC ₂ H ₄ O) ₃ (CH ₃ O) ₃ TP	108
Figure 4.2.16. TGA curves of both chromonic materials	111
Figure 4.2.17. DSC trace of (GlcOC ₂ H ₄ O) ₆ TP and asym-(GlcOC ₂ H ₄ O) ₃ (CH ₃ O) ₃ TP	112

Figure 4.2.19. Textures of chromonics at room temperature from isotropic phase	113
Figure 4.2.20. Textures of $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$ at 175°C from isotropic phase	113
Figure 4.2.21. Columnar phase traped for $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$ at room temperature	114
Figure 4.2.22. Lyotropic phase behaviour for both chromonic materials	114
Figure 4.2.23. Optical texture of $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$ in the columnar phase	115
Figure 4.2.24. Optical texture of hydrated chromonics in the columnar phase	115
Figure 4.2.25. IR spectra for both chromonic materials at 25 °C	116
Figure 4.2.26. A typical UV-Vis spectrum of chromonic materials	116
Figure 4.2.27. Concentration dependent ^1H NMR chemical shifts, recorded in D_2O ($\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$, 25 °C, 400 MHz)	118
Figure 4.2.28. Self-assembly of $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$ based on ^1H NMR measurements	118
Figure 4.2.29. Temperature dependent ^1H NMR chemical shifts, recorded in D_2O ($\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$, 400 MHz	119
Figure 4.2.30. Temperature dependent ^1H NMR of $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$ at 25.1×10^{-3} M and asym- $(\text{GlcOC}_2\text{H}_4\text{O})_3(\text{CH}_3\text{O})_3\text{TP}$ at 12.3×10^{-3} M recorded in D_2O	119
Figure 4.2.31. ATR spectrum of $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$ in water at 25 °C	121
Figure 4.2.32. ATR spectrum of variety of concentration of $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$ at hydroxyl vibration of LCLCs , asym- $(\text{GlcOC}_2\text{H}_4\text{O})_3(\text{CH}_3\text{O})_3\text{TP}$ shows similar behaviour	121

List of Tables

Chapter 2

Table 2.1.	Classification of surfactants based on HLB	21
------------	--	----

Chapter 4

Table 4.1.1.	Influence of the source of Cu(I) on synthesis yields	81
Table 4.1.2.	Effect of temperature on click chemistry	82
Table 4.1.3.	Naming conventions of the β -anomer ATGs products	84
Table 4.1.4.	^1H NMR signals for 4-(2,3,4,6-tetra- <i>O</i> -acetyl- β -D-glucopyranosyl-oxymethyl)-1-dodecyl-1,2,3-triazole	85
Table 4.1.5.	Naming conventions of the α -anomer ATGs products	87
Table 4.1.6.	^1H NMR signals for 4-(β -D-glucopyranosyl-oxymethyl)-1-dodecyl-1,2,3-triazole and 4-(α -D-glucopyranosyl-oxymethyl)-1-dodecyl-1,2,3-triazole	88
Table 4.1.7.	Solubility of the prepared ATGs	89
Table 4.1.8.	Experimental data of the prepared ATGs	91
Table 4.1.9.	Comparison CMC values between glycolipide and ATGs	92
Table 4.1.10.	Summary of phase sequence by using OPM	94
Table 4.1.11.	Experimental data of ATGs by using DSC and OPM	96
Table 4.1.12.	Lyotropic phase behaviour of all ATGs products	98
Table 4.2.1.	^1H NMR assignment of signals for $(\text{GlcOC}_2\text{H}_4\text{O})_6\text{TP}$	109
Table 4.2.2.	^1H NMR assignment of signals for asym- $(\text{GlcOC}_2\text{H}_4\text{O})_3(\text{CH}_3\text{O})_3\text{TP}$	110
Table 4.2.3.	Results of chromonics data by DSC	112
Table 4.2.4.	UV-Vis results of chromonic materials	117

List of Abbreviations

2-D	2-dimensional (NMR)
Ac	Acetyl
Ac ₂ O	Acetic anhydride
APG	Alkylpolyglucoside
Ar	Aryl
ATG	Alkyl triazole glycoside
BF ₃ .Et ₂ O	Boron trifluoride diethyl etherate
br	(NMR) broad signal
CD ₃ OD	Methanol-d ₄
CDCl ₃	Chloroform-d
CH ₂	Methylene
CH ₃	Methyl group
CMC	Critical micelle concentration
Col	Columnar
Cr	Crystals
Cub, Q	Cubic
d	(NMR) doublet
D ₂ O	Deuterium oxide
DCM	Dichloromethane
dd	(NMR) doublet of a doublet (double doublet)
ddd	(NMR) doublet of doublet of doublet
DEPT	Distortionless Enhancement by Polarization Transfer
DFT	Density Functional Theory
DMF	N, N-dimethylformamide
DMSO	Dimethyl Sulfoxide
DMSO-d ₆	Dimethyl sulfoxide-d ₆
DSC	Differential scanning calorimetry
e.g.	for example
EDTA	Ethylenediaminetetraacetic acid
et al.	and others
etc.	and the others
EtOAc	Ethyl acetate
EtOH	Ethanol
Glc	Glucose
h	hour(s)
H ₁	Normal hexagonal
HMQC	C-H Correlation Spectroscopy
HPLC	high performance liquid chromatography
I	Isotropic
IR	Infra-Red spectroscopy
IUPAC	International Union Pure and Applied Chemistry
J	coupling constant
LAS	linear alkyl benzene sulfonates
LC	Liquid crystal

LCLC	Lyotropic chromonic liquic crystal
L_a	Lamellar phase
M	(NMR) multiplet
m_c	Multiplet center
Me	Methyl
MeOH	Methanol
min	Minute(s)
N_2	Nitrogen gas
NaOAc	Sodium acetate
NaOMe	Sodium methoxide
NMR	Nuclear Magnetic Resonance
Nu	Nucleophile
OPM	Optical Polarizing Microscope
POE	Polyethoxyethylene
p-TsOH	para-Toluenesulfonic acid
Py	Pyridine
ROH	Alcohol
Rt	Room temperature
s	singlet
$SnCl_4$	Tin tetrachloride
T	(NMR) triplet
THF	Tetrahydrofuran
T_K	Krafft temperature
TLC	thin layer chromatography
TP	Triphenylene core
Ts	Toluenesulfonyl
Γ	Surface tension